

## Chapter 11 Experiment 17 Precipitation Reactions

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Experiment 17 Precipitation Reactions 109 Name Date Class Set 3. 0.1M barium chloride, BaCl2. 0.1M magnesium nitrate, Mg(NO3)2. 0.2M sodium chromate, Na2CrO4. 0.2M aluminum sulfate, Al2(SO4)3. 0.2M potassium chromate, K2CrO4. 0.2M silver nitrate, AgNO3. SAFETY FIRST! In this lab, the solutions you use may contain harmful materials.

**Name**

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The aim of this part of the experiment is to determine the factors that control whether or not a precipitate forms, i.e. is it the sum of the ion concentrations (in this case [Ag + ] + [Cl - ]) or the product of the ion concentrations (in this

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Aqueous silver nitrate (AgNO 3) is added to a solution containing potassium chloride (KCl), and the precipitation of a white solid, silver chloride (AgCl), is observed: AgNO 3 (aq) + KCl (aq) → AgCl (s) + KNO 3 (aq) Note that the product silver chloride is the precipitate, and it is designated as a solid.

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This chapter provides an overview of downstream processing. After completing this chapter, students will be able to: describe the general elements of the downstream purification process for a typical monoclonal antibody-based therapeutic, starting with the bioreactor harvest pool and finishing with pre-formulation, pure drug product.

~~Chapter 11 Downstream Processing - Biomanufacturing~~

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