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ox:  $[Ag \rightarrow Ag^+ + e^-] \times 2$ ; red:  $Cu^{2+} + 2e^- \rightarrow Cu$ . balanced:  $2Ag(s) + Cu^{2+}(aq) \rightarrow 2Ag^+(aq) + Cu(s)$   $E^\circ = E^\circ_{ox}(Ag) + E^\circ_{red}(Cu^{2+}) = 0.80V + 0.34V = 0.46V$ . Not spontaneous. ox:  $[Ni(s) \rightarrow Ni^{2+} + 2e^-] \times 5$ ; red:  $[5e^- + 8H^+ + MnO_4^- \rightarrow Mn^{2+} + 4H_2O] \times 2$ .

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